

The criss-cross method of balancing charge!

Ionic compound formulas must contain the fewest number of ions that “balance” out positive and negative charge (the same amount of each). The “criss-cross” method is one way of writing the formulas properly. The formula for ionic compounds is called a “formula unit.”

1. Write symbols and charges of ions.

2. Crisscross:

The cation charge becomes the anion subscript

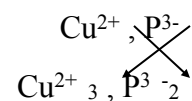
The anion charge becomes the cation subscript

3. Clean up format for final answer

- do NOT write ionic charges

- reduce subscripts to lowest ratio

- do NOT write the subscript “1”



Use the criss-cross method to write formula units for these ionic compounds
(2 examples are done for you)

	Cl^-	O^{2-}	P^{3-}	S^{2-}
Na^+	NaCl	Na_2O		
K^+				
Ba^{2+}				
Fe^{2+}				
Cr^{3+}				
Li^+				
Mg^{2+}				
Al^{3+}				
Ga^{3+}				
Sn^{2+}				
Ca^{2+}				